Isothermal Vapor-Liquid Equilibria for the binary system of carbon dioxide+ iodomethane at five temperatures from 283.15 to 323.15 K

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Isothermal Vapor-Liquid equilibria data for the binary system of carbon dioxide+ iodomethane were measured at five temperatures within the range 283.15-323.15K. The measurements were carried out by using a circulation-type equilibrium apparatus in which both vapor and liquid phases were recirculated. The experimental data were correlated the PR-EoS using the Wong-Sandler mixing rules combined with the NRTL excess Gibbs free energy model and the PR-EoS using the Universal mixing rule. It is confirmed that the data calculated by this equation of state is in good agreement with experimental data.